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This classical expression for the second virial coefficient was first derived by Leonard Ornstein in his 1908 *The Virial Coefficients of Pure Gases and Critical constants and second virial coefficients of gases* 3.5

where R is the gas constant: $R = (8.314 510 000 070) \text{ J K}^{-1} \text{ mol}^{-1}$; B, C, \dots are the second, third, ... virial coefficients which for a given substance are functions of temperature only; and Z is called the compression factor.

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This critical compilation of virial coefficients of pure gases is a sequel to *The Virial Coefficients of Pure Gases and Mixtures*, by J.H. Dymond and E.B. Smith (Oxford University Press, 1979).

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This paper describes the first part of an investigation of the thermodynamic properties of gases and gaseous mixtures undertaken a few years ago at the National

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The deviation is best stated in a virial expansion of molar A is the first virial coefficient. There is a large gap between pure gaseous phase and pure

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The Pitzer equation for the calculation of the second virial coefficients of pure compounds, written in its more updated form according to Tsonopoulos

Cross Second Virial Coefficients and Dilute Gas Transport ...

pure gases. All PESs are based data for the second virial and cross second virial coefficients. Overall, the agreement of the calculated thermophysical property