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So far the role of the Van der Waals effect has been that of a nasty, disturbing phenomenon, something to be eliminated at all costs. But it is precisely in this latter respect where almost all solvent effect studies fall short. Not only is σ_w usually large (larger than σ_a and σ_e even in ^1H NMR and probably the dominating term with heavier nuclei), but it is strongly variable from one

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Van der Waals Forces in NMR Intermolecular Shielding Effects Frans H. A. Rummens Department of Chemistry University of Regina Regina, Saskatchewan S4S 0A2, Canada

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The usefulness of solvent effect studies on NMR chemical shifts need not be elaborated here; many applications of solvent effects continue to be published in great profusion. Quite a few intermolecular phenomena may contribute to solvent shifts, but there is always the ubiquitous Van der Waals

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Van der Waals Forces and Shielding Effects (NMR Basic Principles and Progress, Band 10) (Englisch) Gebundenes Buch 1. Dezember 1975

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The usefulness of solvent effect studies on NMR chemical shifts need not be elaborated here; Contrary to such other effects as neighbour anisotropy σ_a , reaction field contribution σ_e or complexation effects σ_e , no major direct use has yet been found for the Van der Waals effect.

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Van der Waals forces' is a general term used to define the attraction of intermolecular forces between molecules. There are two kinds of Van der Waals forces: weak London Dispersion Forces and stronger dipole-dipole forces .

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